Aug 20th, 3:15 PM - 3:45 PM

Hydrogen Fuel Cells and Storage Technology (FCAST) Project

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Hydrogen Fuel Cells and Storage Technology (FCAST) Project

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UNLV Research Foundation

sponsored by the DOE EERE Hydrogen Research Program through the UNLV Research Foundation, contract number DE-FG36-05GO85028

This presentation does not contain any proprietary, confidential, or otherwise restricted information
FCAST partners

UNLV Experiment

• Chulsung Bae – Chemistry
• Andrew Cornelius – Physics
• B.J. Das – Electrical Engineering
• David Hatchett – Chemistry
• Clemens Heske – Chemistry
• Wayne Stolte, Oliver Hemmers, Dennis Lindle – Chemistry

UNLV Theory

• Changfeng Chen – Physics
• Eunja Kim – Physics
• Steven Lepp – Physics
• Bala Naduvalath – Chemistry
• Tao Pang – Physics
• Bernard Zygelman – Physics

External Partners

• United Technologies (UTC) Power
• Rice University
• Lawrence Berkeley National Lab
• Air Products
• Hahn-Meitner-Institute, Berlin
• Shanghai Jiatong University
• Penn State
Why Hydrogen Economy?
Why Hydrogen Economy?

- We need a new fuel! (sooner or later)
Why Hydrogen Economy?

- We need a new fuel! (sooner or later)
  - This could be:
    - Biodiesel
    - Hydrogen
    - Electricity
    - Uranium
    - ...

What would a Hydrogen Economy need?

- Hydrogen Production
- Hydrogen Storage
- Hydrogen Delivery
- Hydrogen Consumption
What would a Hydrogen Economy need?

- **Hydrogen Production**
  - Solar (thermal, photoelectrochemical)
  - Nuclear
  - Currently: natural gas reforming

- **Hydrogen Storage**
  - Nanomaterials, Metal Hydrides, Chemical Hydrides

- **Hydrogen Delivery**
  - Pipelines, Trucks, Tanks, ...

- **Hydrogen Consumption**
  - Fuel Cells, Internal Combustion Engine
What would a Hydrogen Economy need?

• Hydrogen Production
  • Solar (thermal, photoelectrochemical)
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• Hydrogen Delivery
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• Hydrogen Consumption
  • Fuel Cells, Internal Combustion Engine
Objectives of FCAST

• Perform closely-coupled theoretical and experimental investigations of
  – hydrogen adsorption/desorption in various matrices to establish a solid understanding of optimal storage concepts
  – the electronic and geometric structure of metal hydrides, nanomaterials (C, B, N, transition metals, alloys), metal adatoms, and adsorbed hydrogen molecules/atoms
  – Fuel cell membranes and catalytic materials to predict optimized materials and structures for hydrogen storage and fuel cells in the DOE Hydrogen program

• Collaborate closely with external partners
Approach

Task 1: Theory and Experiment of Nanomaterials for Storage Applications
(New Materials, Hydrogen Uptake, Local Electronic Structure, Adsorption Energies and Geometries, …)

Task 2: Metal Hydrides (Structure, Reversibility, T- and P-Dependence, …)

Task 3: Mesoporous Polymer Nanostructures (Synthesis, Hydrogen Uptake, …)

Task 4: Improved Fuel Cell Membrane

Task 5: Design and Characterization of Improved Fuel Cell Catalytic Materials
Titanium-decorated Carbon Nanotubes as a Potential High-capacity Hydrogen Storage Medium

a  C₈TiH₈ (5.3 wt%)  b  C₄TiH₈ (7.7 wt%)

FIG. 3 (color online). Two high-density hydrogen coverage on a Ti-coated (8, 0) nanotube.
Single-Walled Carbon NanoTubes (SWNT)

- nanotubes come in bundles
- Samples have a size distribution
- Samples contain impurities

A local experiment on one individual nanotube would be ideal!

TEM of SWNT

Which is the best nanotube for hydrogen storage?
How does hydrogen adsorption/desorption work?
How can we improve it?
• Up to 4 H₂ are adsorbed on each Ti atom with the binding energy ranging from 0.1 eV to 0.4 eV per H₂. (7.8wt% for double side coverage)

The binding energy of H₂ on Sc is slightly lower than that on Ti.
A novel class of 3D nanoframeworks based on CNTs (Balakrishnan Naduvalath, Chemistry, Eunja Kim, Physics, Task 1)

- Car-Parrinello molecular dynamics simulations indicate that the proposed frameworks are thermodynamically stable up to 20 ps at 300 K and 2 ps at 600K
- Preliminary results indicate that Li-decorated 3D nano-frameworks are promising for hydrogen storage
Ti₅
1.8694 eV

Ti₇
2.3128 eV

Ti₈
2.3562 eV

Ti₉
2.4472 eV

Ti₁₀
2.5187 eV

Ti₁₁
2.5700 eV

Ti₁₃
2.8063 eV

Ti₁₅
2.9270 eV

Electronic structure of Titanium clusters (B. Naduvalath, Task 1)

- Tiₙ clusters evolve on Pentagonal growth pattern
- Second energy difference indicates Ti₇ and Ti₁₃ clusters are highly stable, which agrees well with the experimental results

Second Energy Difference

\[ \Delta_2 E = E(n+1) + E(n-1) - 2E(n) \]
**Ti$_{13}$ cluster and H$_2$ saturation (B. Naduvalath, Task 1)**

- Hydrogen multi-center bonds in Ti$_{13}$H$_m$
- $\mu_3$ for $m \leq 20$ and $\mu_2$ in Ti$_{13}$H$_{30}$
- Cage expansion due to saturation from $m = 20 – 30$ by 6%
Surface and interface spectroscopy/microscopy of nanomaterials for hydrogen storage (Clemens Heske, Chemistry, Task 1)

Experiment matrix for Hydrogen storage on (metal-decorated) carbon nanomaterials:
- Carbon (nano)materials: \( \text{C}_6\text{O}, \text{SWNT}, \text{HOPG} \)
- Metal (co-)adsorbates: Ti, Li
- Hydrogenation: molecular, atomic

- How does Hydrogen interact with carbon-based nanomaterials?
- Why is there a “gap” between theory and experiment?
Soft X-ray spectroscopies

- Photoelectron-Spectroscopy (PES)
- Auger-Electron-Spectroscopy (AES)
- X-Ray Emission Spectroscopy (XES)
- X-ray Absorption Spectroscopy (XAS)
High dynamic range
XPS, UPS, Auger, IPES

High resolution
XPS, UPS, Auger

Glovebox

Scanning Probe Microscope

Sample preparation and distribution
Scanning Tunneling Microscopy (STM) / Spectroscopy (STS)

**Microscopy**
- Tunneling tip can be scanned over the sample by piezos
- Tunneling current is measured at a tunneling voltage $V$ and kept constant by a feedback controlling the z-direction → “topography of electron density”

**Spectroscopy**
- Tunneling voltage $V$ is varied at one spot with constant tip-sample distance → density of states around $E_F$

$$I(V) = \int_0^{eV} \rho_s(E) \cdot \rho_t(E - eV) \cdot T(E, eV, d) dE$$

$$\rho_s(eV) \approx c \cdot \frac{dI(V)}{dV}$$
STM on SWCNT – 2

- Electro-deposited (6 min, **upside down**) on HOPG, SWCNT (not cut)

- Electro deposition worked
- Low coverage with long SWCNT
- “Small” bundles (e.g. two bundled tubes) and individual tubes are observed
- Atomic resolution can be achieved
Scanning Tunneling Microscopy/Spectroscopy of SWNT with/without Ti decoration (Task 1)

STM image of SWNT on Au with atomic resolution

I-V curve and STS of SWNT on Au
Atomic Hydrogen Source (AHS)


Atomic hydrogen is produced by high temperature thermal cracking of molecular hydrogen via a W capillary tube, submitted to HV and electron bombardment

Atomic Hydrogen Source components:

• Gas Reservoir

• Tungsten Capillary Tube
  – 50-mm long, 1.6-mm O.D, 0.6-mm I.D.
  – Acceleration Voltage of ~+3 kV

• Tungsten Filament
  – 1.7 A, 7.0 V, 12 W
Atomic/molecular hydrogen ratio

$P_{\text{AT}}/P_{\text{H}_2}$

W Capillary tube Temperature (K)

9.0x$10^{-8}$ mBar

PAH/PH2
1. No shift in C 1s for molecular hydrogen adsorption (at RT)

2. C 1s shifts to higher binding energy for atomic hydrogen (along with capillary temperature), indicating H adsorption
XPS: Hydrogenation of SWNT (Task 1)
Gibbs free energy and temperature-pressure phase diagram of lithium alanates (Changfeng Chen, Task 2)

Apply first-principles electronic structure and lattice dynamics calculations within and beyond the harmonic phonon approximation to examine the thermodynamic phase stability of lithium alanates and predict their reaction pathways and reversibility

**Results:**

- Obtained a comprehensive set of thermodynamic functions over a wide temperature range for LiAlH$_4$, Li$_3$AlH$_6$ and LiH.
- Evaluated decomposition reactions to determine reversibility and suitability for practical use in mobile applications.
- Established the thermodynamic (temperature-pressure) phase diagram for lithium alanates and identified key operating physical parameters for hydrogen storage and reversible release-recharge process.
Pd(ii) reduction in PANI

Pd morphology is a function of the number of voltammetric cycles.

Pd aggregation also possible with potentiometric growth.
PANI/Pd Composites (Task 3)

Pd(iv) Reduction in PANI

Pd thickness is a function of the number of voltammetric cycles
H Sorption Apparatus (Task 3)
Hydrogen Sorption in Chemical Composites (Task 3)

Material A

Sorption is obtained using a normal GC with a hydrogen generator

- The first peak remains unchanged relative to the second because it represents the void volume of hydrogen in the tube rather than sorbed hydrogen
- The second peak represents sorbed hydrogen
- A temperature ramp is used to observe desorption
Hydrogen Sorption Results and Conclusions (Task 3)

- Five composite materials have been produced that show promise for Hydrogen sorption.
- Preliminary measurements have been made to verify the sorption properties.
- Variations in the chemical composites have been eliminated by treatment with $\text{NaBH}_4$ thus reducing any unreduced species.
- This material shows the highest sorption suggesting that treatment of the other chemically prepared composites may increase sorption properties.
PEM Fuel Cell
Summary

FCAST is a joint experimental and theoretical project to enhance the understanding of hydrogen fuel cells and storage materials

- Joint experimental and theoretical work performed on electronic structure of carbon nanoclusters
- Stable structures of graphitic-BC$_2$N as potential hydrogen storage media identified
- The electronic structure of Ti decorated SWCNTs explored using X-ray and electron spectroscopy. Significant oxidation of Ti leading to TiO$_2$ formation is observed
- Systematically explored hydrogen uptake of transition metal-bonded organometallic systems (Sc, Ti, V) using DFT methods
- Proposed new class of carbon nanoframeworks (thin SWCNTs linked by phenyl spacers) as potential hydrogen storage media
- Investigated electronic structures and bondings in hydrogen saturated Ti and Ti-Al clusters and identified novel bonding motifs which may be harnessed to design novel hydrogen storage systems
- Synthesized bulk quantities of mesoporous PANI/Pd composites for hydrogen storage

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